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of the new solution? (Answer: 9-801 HN03) =  $M_1V_1 = M_2V_2$  (  $0.8M \cdot 30.00 \text{ ml} = M_2 \cdot 50.00 \text{ ml}$ , of water is added to 20.00 ml, of 0.0500 M sodium carbonate. What is the molarity of sodium carbonate in the new solution? (Answer: 0.0200 M) (20.00 ml C (350.00 casF) 4. What volume of 0.100 M NaCl is needed to make 5.00 ml, of 0.200 M solution? (Answer: 0.01 (0.100 M) V)

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$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} = M = \frac{\text{mol}}{\text{L}}$$
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