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Note : # Neutrons = Mass Number –

Atomic Number. # Neutrons of

Nitrogen-14 = $14 - 7 = 7$. # Neutrons of

Nitrogen-13 = $13 - 7 = 6$. c) The number

of electrons in a neutral atom always

equals the atomic number. The atomic

number of both isotopes is 7, therefore the

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number of electrons in both isotopes is also 7.

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